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COD

fTPHĚ

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ly /y :

ly /y :

fTPHĚ

fPAH_sĚ

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f_v ççĚ

UV

pH_iH₂O₂

éL

pH_iH₂O₂

COD

pH=é ÿ / M

ÿ / mM

COD ñ / ž

UV

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pH"

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UV

COD

fPH= Ě

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UV/Fe²⁺/H₂O₂

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fTPHĚ

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در این مطالعه، فرایند فتوفتوتون (FOT) به عنوان یک روش نوین برای حذف آلاینده‌ها از آب و خاک مورد بررسی قرار گرفت. این فرایند با تابش اشعه فرابنفش (UV) همراه با وجود پراکسید هیدروژن (H_2O_2)، پراکسیل‌ها ($\cdot OH$) را تولید می‌کند که قادر به اکسیداسیون و تجزیه آلاینده‌ها هستند. فرایند FOT در مقایسه با سایر روش‌ها، مزایایی نظیر عدم تولید اسید سولفوریک، عدم نیاز به مواد شیمیایی جانبی و امکان حذف آلاینده‌های难降解 (refractory) را دارد.

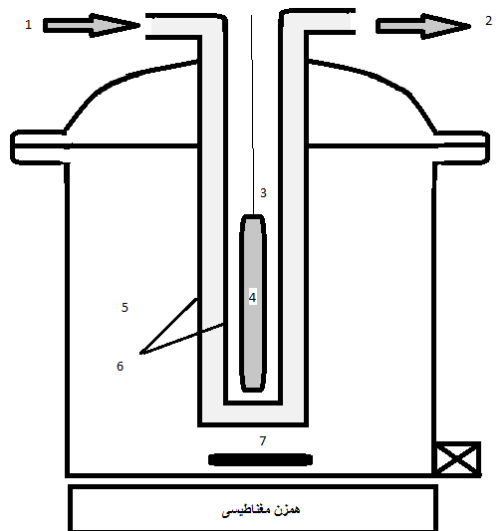
در این مطالعه، اثر پارامترهای مختلف بر کارایی فرایند FOT در حذف ماده آلاینده PAH₅ بررسی شد. پارامترهای مورد بررسی شامل دوز اشعه UV، دوز H_2O_2 ، دوز ماده آلاینده، pH و دما بودند. نتایج نشان داد که افزایش دوز اشعه UV و دوز H_2O_2 منجر به افزایش کارایی حذف آلاینده می‌گردد. همچنین، دمای بالاتر و pH کمتر منجر به بهبود کارایی فرایند شده است.

برای مقایسه، کارایی فرایند FOT با روش‌های سنتزیکال و بیولوژیکی (Bioremediation) مقایسه شد. نتایج نشان داد که فرایند FOT در حذف آلاینده‌های难降解، کارایی بالاتری در مقایسه با سایر روش‌ها دارد.

در ادامه، مکانیسم واکنش فرایند FOT و معادله‌های سینتیکی آن بررسی شد. برای تعیین ترتیب واکنش، داده‌های تجربی با معادله‌های سینتیکی مقایسه شدند. نتایج نشان داد که واکنش‌ها دارای ترتیب واکنش اول و دوم می‌باشد. همچنین، انرژی فعال‌سازی (E°) برای واکنش‌ها محاسبه شد و در محدوده ۱۰-۱۵ کیلوژول بر مول قرار گرفت.

نتیجه‌گیری: فرایند فتوفتوتون (FOT) یک روش موثر و اقتصادی برای حذف آلاینده‌ها از آب و خاک است. این فرایند با تنظیم پارامترهای مختلف، کارایی خود را بهبود می‌بخشد و می‌تواند به عنوان یک روش جایگزین در صنایع مختلف مورد استفاده قرار گیرد.

۳۰ min
 pH
 ۰.۱ M NaOH
 pH
 H₂O₂
 pH
 COD
 COD
 pH
 H₂O₂
 pH
 °C
 pH < 7
 DR5000
 COD
 TPH
 HACH
 COD
 COD
 HACH
 COD
 mg/L
 DR5000
 Excel
 COD
 mg/L
 COD₀
 pH
 M
 COD
 mg/L



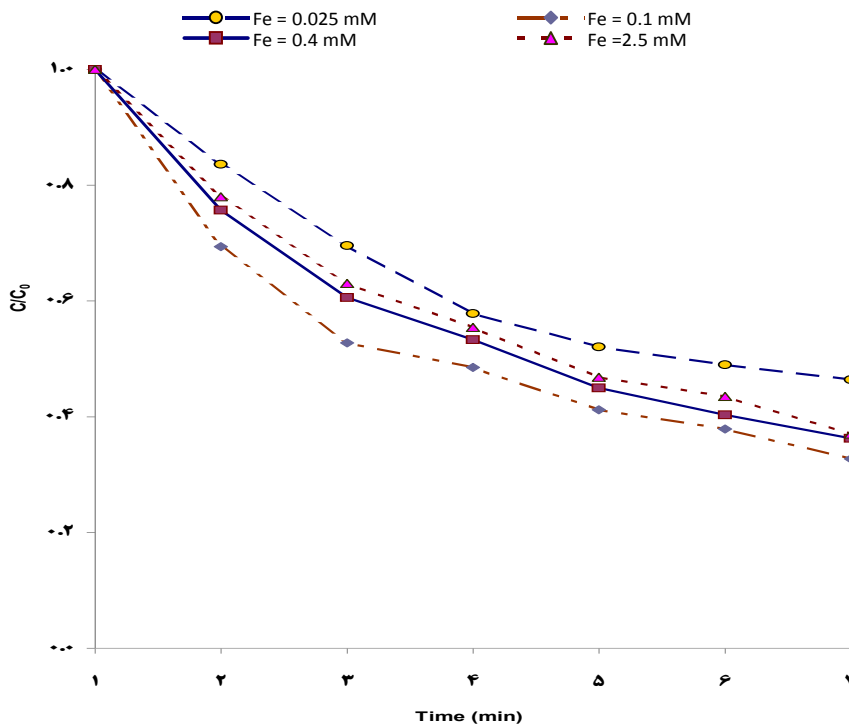
UV
 pH
 COD
 mg/L
 COD₀
 pH
 M
 COD
 mg/L

UV/Fe/H₂O₂
 mM FeSO₄·7H₂O
 pH= M H₂O₂

UV H₂O₂ (L

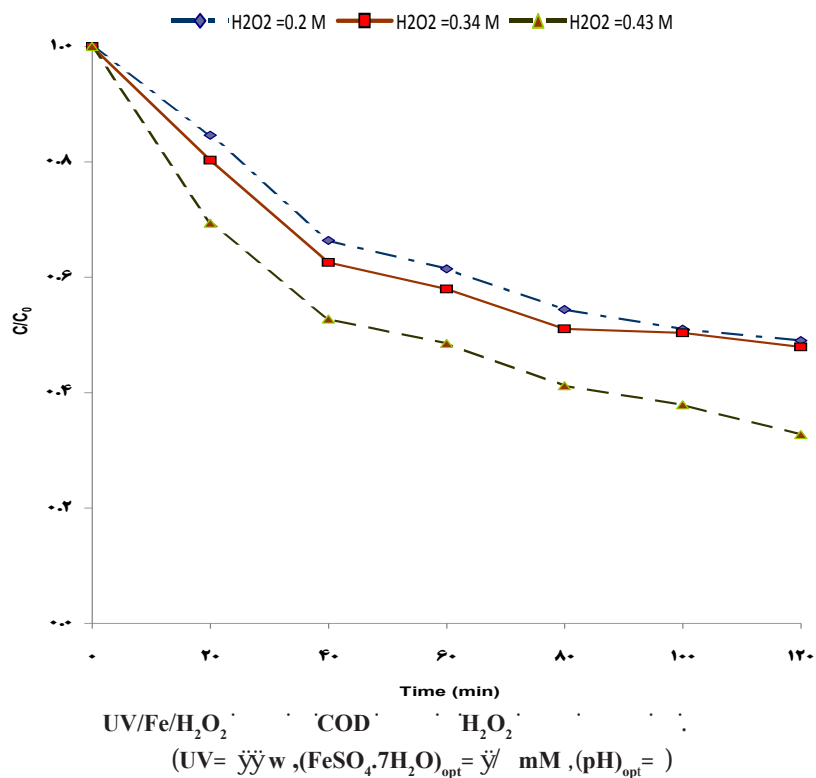
mM COD UV
 UV
 f)

کارایی حذف فرایند UV
 مجزای UV، % بوده است.



UV/Fe/H₂O₂ COD
 (UV= w H₂O₂= mol pH= , COD₀ = - mg/L)

$\dot{y} / m M$ H_2O_2 " " / m M
 \dot{y} / mol " \dot{n}
 \dot{e} pH H_2O_2
 pH " $\dot{y} \cdot \dot{e}$ H_2O_2
 pH μ " UV/Fe/ H_2O_2
 $\dot{y} M_i NaOH$ \dot{y} " "
 UV H_2O_2 " $\dot{y} / m \mu$
 COD " $\dot{y} / \dot{y} / \dot{y} / \mu$ H_2O_2 "
 pH \dot{y} / μ H_2O_2 $\dot{y} / \mu m$ H_2O_2 " $pH =$
 $\dot{n} /$ " $\dot{pH} =$) H_2O_2 " "
 . (\dot{L} " UV " "
 " " "
 $UV/Fe/$ COD " $\dot{y} / m \mu \dot{L}$ COD "
 H_2O_2 $\dot{n} /$ " $H_2O_2 / \dot{y} / mol$ "
 . (\dot{L} " "
 pH "
 min COD " $1/C$ UV/ pH "
 \dot{y} " pH " Fe/H_2O_2



$\ln(C/C_0)$

$n /$

n

UV Fe(II)

$Fe^{+2} + H_2O_2 \rightarrow Fe^{+3} + OH^- + HO^0$

(HO⁰ Fe(II) UV

H_2O_2 Fe(II) Fe(II)

COD Fe(II)

H_2O_2

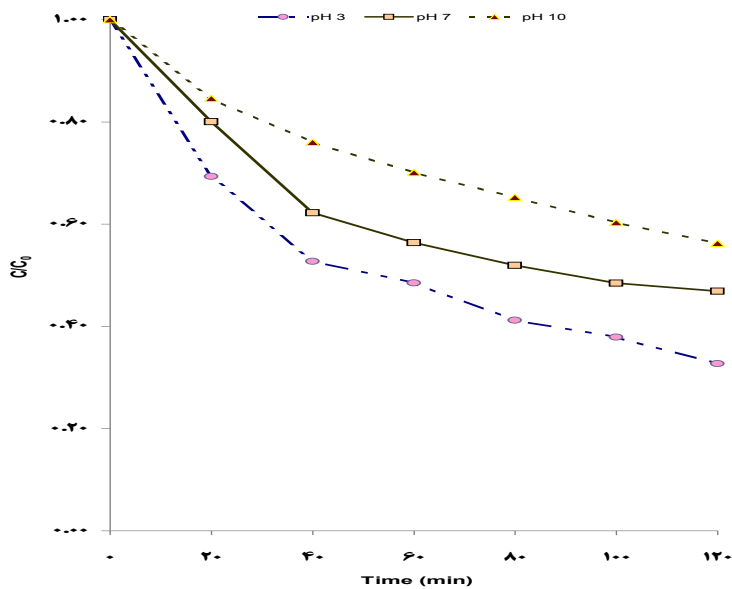
H_2O_2 COD

H_2O_2 mM COD

H_2O_2

H_2O_2

(mM



UV/Fe/H₂O₂ COD pH

(UV= 100 W, (FeSO₄·7H₂O)_{opt} = 1 mM, (H₂O₂)_{opt} = 1 mol)

h BTX

$\dot{y} \text{ mM } i \text{ Fe}^{2+}$ $\dot{y} \text{ mM } i \text{ H}_2\text{O}_2$ $\text{HO}^0 + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{HO}^0_2$

Oswaldo "fl L ñ $2\text{HO}^0_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}^0_2$

Chiavone-Filho

$\dot{y} \text{ mM}$ H_2O_2 " " "

$\dot{y} \text{ mM}$ $-\dot{y} \text{ mM}$ " " "

$\dot{y} \text{ mM}$ " " "

ñ COD ñ / " "

pH "fl L " " "

pH " " " COD :

! " " " " "

pH " " " COD :

UV " " "

! " " " " "

$\dot{y} \text{ mM}$ Kang Hua " "

H_2O_2 pH " " "

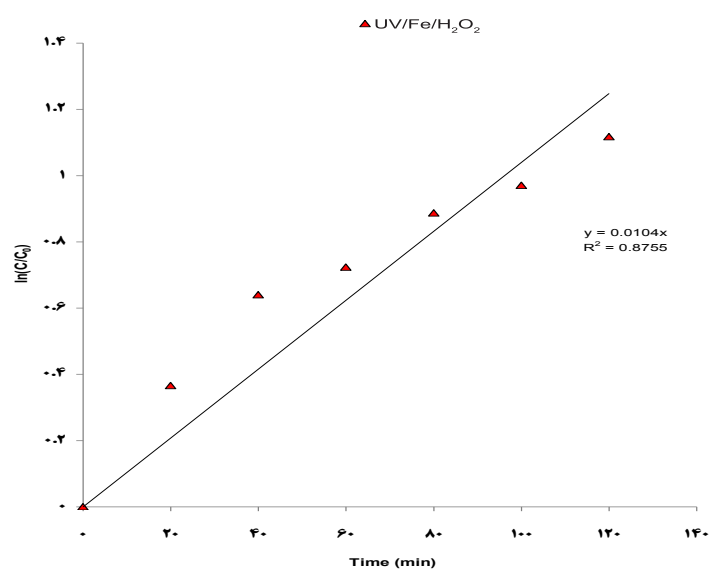
" H_2O_2 pH= \dot{y} " " "

$\dot{y} \text{ mM}$ Raquel F. PupoNogueira " " " BTX

$\dot{y} \text{ mM } i \text{ Fe}^{2+}$ " " "

TritonX-100 (TX- $\dot{y} \text{ mM}$ " " "

$\dot{y} \text{ mM } i \text{ H}_2\text{O}_2$ " " "

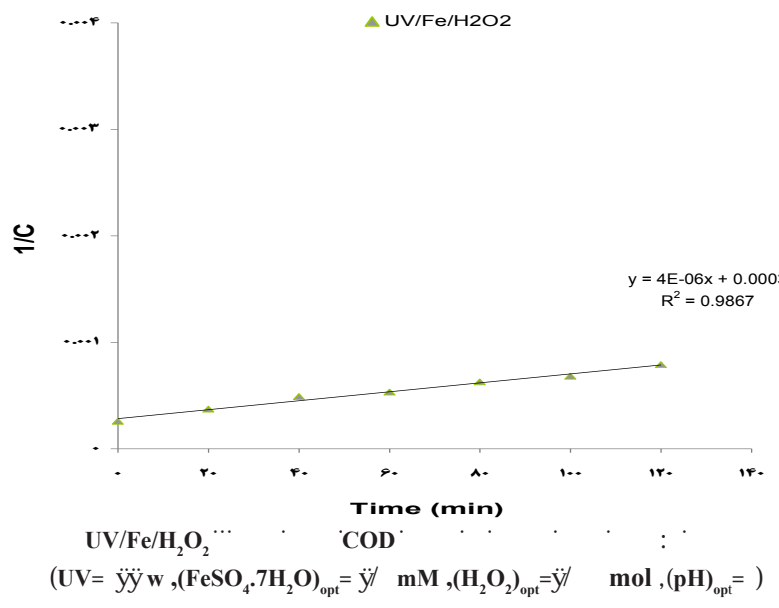


UV/Fe/H₂O₂ COD

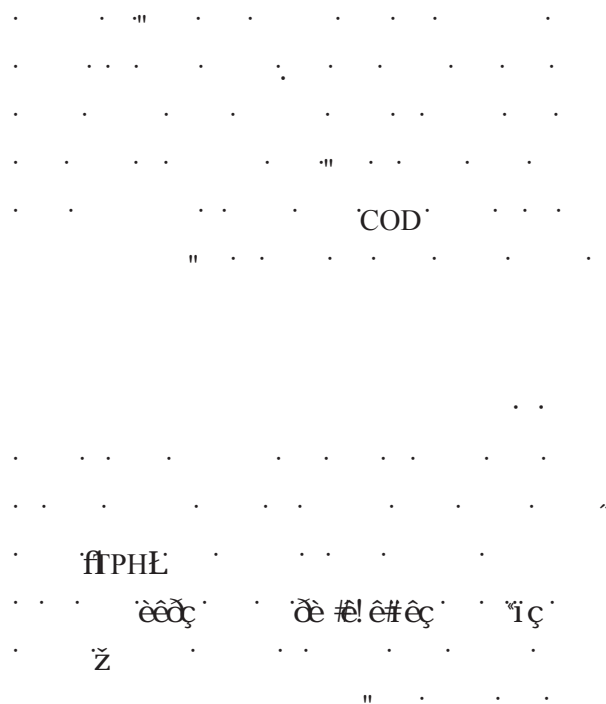
(UV= $\dot{y} \text{ w}$, (FeSO₄.7H₂O)_{opt}= $\dot{y} \text{ m}\mu$, (H₂O₂)_{opt}= $\dot{y} \text{ mol}$, (pH)_{opt}=)

$\frac{1}{C} - \frac{1}{C_0} = k.t$
 COD
 Farrokhi
 UV
 COD
 $pH = 7$

pH
 PupoNogueira
 UV
 J. Watts
 UV/Fe/H₂O₂
 Kavitha Palanivelu
 Farrokhi
 UV
 COD



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Application of Photo-Fenton Process for COD Removal from Wastewater Produced from Surfactant-Washed Oil-Contaminated (TPH) Soils

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ABSTRACT

Background and Objective: The base structure of total petroleum hydrocarbons (TPH) is made of hydrogen and carbon. Widespread use, improper disposal and accidental spills of this compounds lead to long term remaining of contaminations such as organic solvents and poly aromatic hydrocarbons (PAHs) in the soil and groundwater resources, resulting in critical environmental issues. In this study, an oil-contaminated soil was washed using Tween 80 surfactant and the application of photo-Fenton process (UV/Fe²⁺/H₂O₂) for treatment of the produced wastewater was evaluated.

Materials and Methods: Tween 80 is a yellow liquid with high viscosity and soluble in water. In order to determine of the photo-Fenton process efficiency, we studied effective variables including Fe concentration, pH, H₂O₂ concentration, and irradiation time. The UV irradiation source was a medium-pressure mercury vapor lamp (400 w) vertically immersed in the solution within 2 L volume glass cylindrical reactor.

Results: The results showed that efficiency of COD removal depends on the initial Fe concentration, pH, H₂O₂ concentration and irradiation time.

Under optimum conditions, (Fe: 0.1 mM, H₂O₂: 0.43 mM, pH: 3 and UV light irradiation time: 2 hours) the removal efficiency of COD was 67.3%. pH plays a crucial role in the photo-Fenton process such that the removal efficiency increased with decreasing of pH.

Conclusion: According to the results of this study, under acidic condition, this process is an efficient method for COD removal from the wastewater studied.

Keywords: Total Petroleum Hydrocarbon (TPH), Tween 80, Advanced oxidation, UV/Fe²⁺/H₂O₂ process

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