

ž ž ž ž

ü TiO₂/UV

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ly /y : y / / :

(Ag-TiO₂)

yy yi ymg/L

i pH

/ y / i y# g/L

Ag-TiO₂

Ag-TiO₂

y g/L Ag-TiO₂

pH:

yymg/L

fn / E

/ g/L

y g/L

n /

n

Ag-TiO₂

fl E

! ! ! ! !

UV-E (ZnO-E TiO₂-E) (E) " "

"fl-E "fl-E

TiO₂ " - mg/L

" fl₂nm-E (E) TiO₂ "

Ô "fl-Doping-E (E) "

"fl-E

P25 TiO₂ " Hombikat (E) "fl-E

Ô " TiO₂ " (y-E) "

fl-E "fl-E (Photocatalytic Degradation-E) "

Ag-TiO₂ UV
 pH
 TiO₂ Ag-TiO₂ Degussa ,25
 TiO₂- Ag
 (Ag/Ti)
 Scanning Electron Microscope-Energy Dispersive)
 Seron Technology AIS-2100 (X Ray (SEM-EDX
 X' Pert MPD
 (Transmission
 ZEISS-EM10C Electron Microscopy (TEM)
 KV (Accelerating voltage)
 Brunauer- Emmett- Ag-TiO₂
 Autosorb 1 Quantachrome Teller (BET
 nm
 fl

HCOOH
 Hole Scavenger
 Ag-TiO₂
 (Photodeposition
 Hydrothermal Sol-gel
 Chemical Photoreduction
 Vapor Deposition
 Swamiathan
 Tryba DB53 DR23
 UV
 Shirzad Siboni
 UV/TiO₂
 pH
 pH= min
 mg/L g/L
 Ghanbarian
 UV TiO₂
 LAS

NaOH pH "

" $\dot{y} / \dot{y} \dot{y} \dot{y}$ N · HCl $\dot{y} / \dot{y} \dot{y} \dot{y}$ N / L L

" "

" $\dot{y} \dot{y}$ $\dot{y} \dot{y}$ mg/L cm fl WL

" cm \dot{y} cm

" "

i \dot{y} i \dot{y} i \dot{y} i \dot{y} min

" "

" \dot{y} \dot{y}

" "

" \dot{y} mL

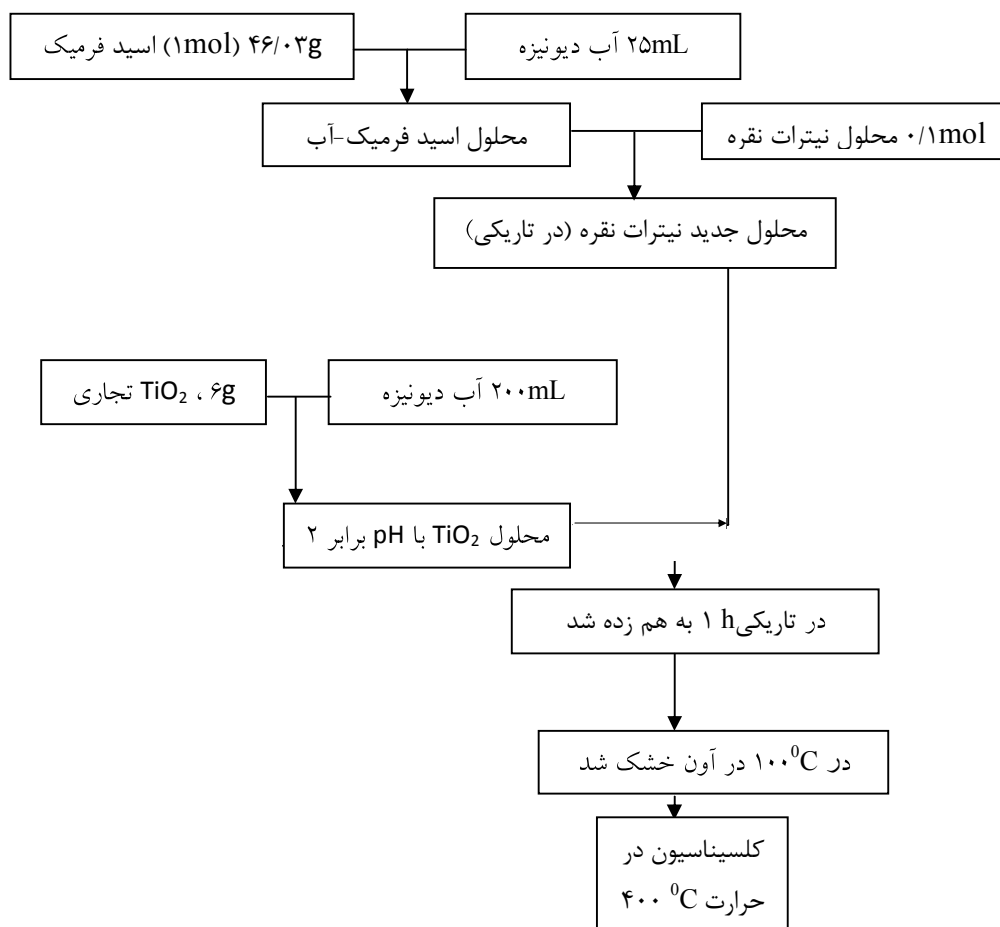
\dot{y} min $\dot{y} \dot{y} \dot{y}$ rpm mg/L

" "

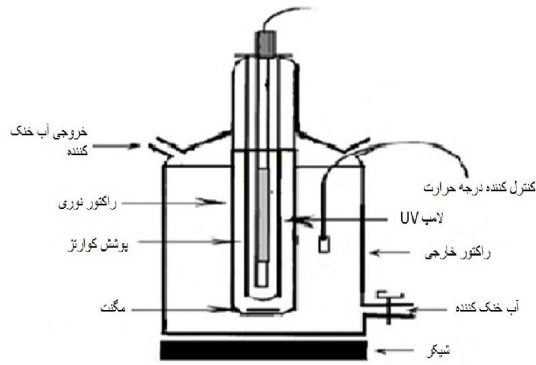
\dot{y} / μ $\dot{y} \dot{y}$ \dot{y} i \dot{y}

Perkin-Elmer Lambda g/L Ag-TiO₂

\dot{y} nm 15-UV/Vis Spectrometer pH= i / \dot{y} i \dot{y} i \dot{y}



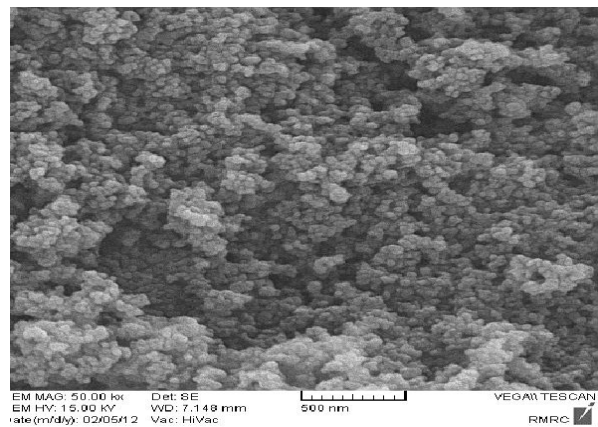
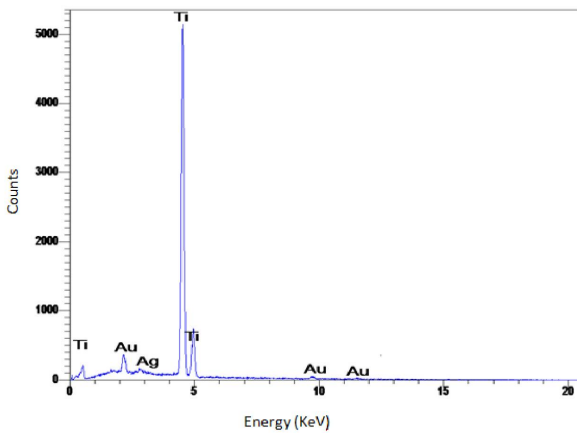
Ag-TiO₂



SEM-EDX
 Ti Ag (at%
 wt%
 Ag/TiO₂
 Au
 TEM SEM
 (/ nm
 EXRD
 (/ Ag-TiO₂
 P25)
 BET (TiO₂ Degussa
 Ag-TiO₂
 TiO₂-P25 doped TiO₂
 Ag-TiO₂ ± m²/g TiO₂- P25
 / m²/g
 pH pH

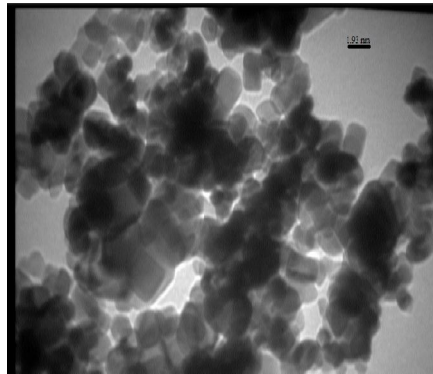
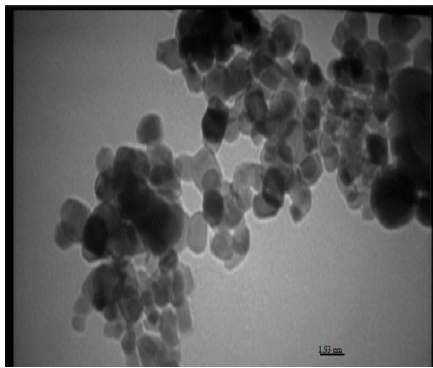
Perkin-Elmer
 Lambda 25-UV/Vis Spectrometer Elmer
 (DR5000 nm
 Ag-TiO₂ UV
 UV
 pH
 Ag-TiO₂/UV
 UV
 SPSS16

Ag-doped TiO₂ SEM-EDX



SEM

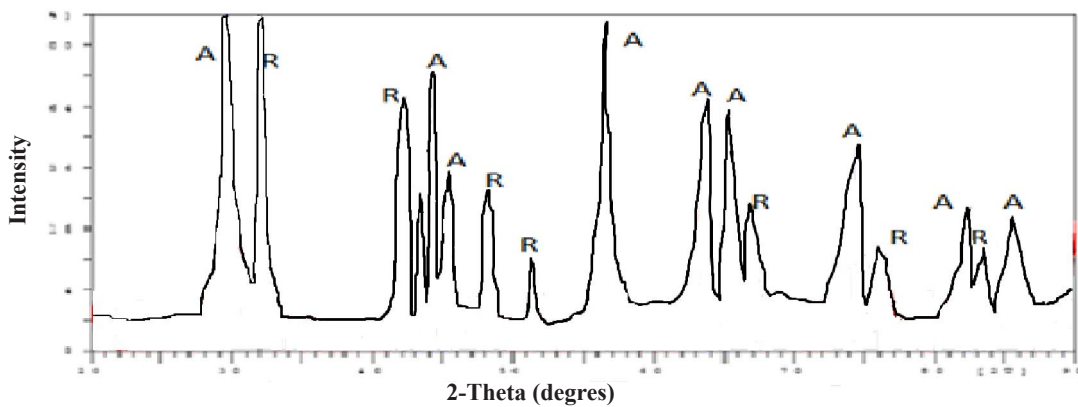
SEM-EDX



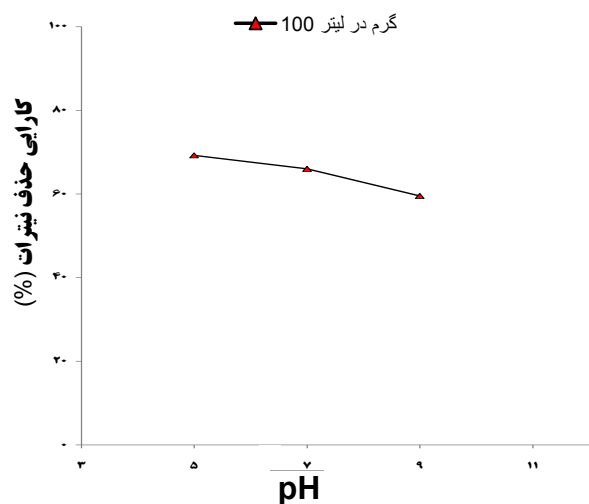
TEM

Ag-TiO₂ / Ag-TiO₂/UV
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7

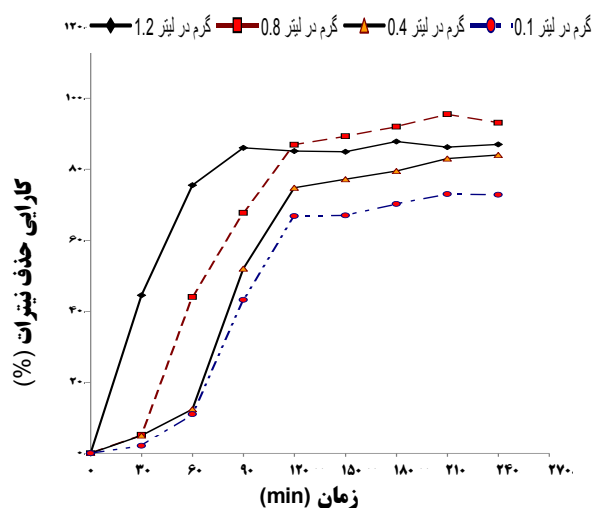
Ag-TiO₂ / Ag-TiO₂/UV
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7
 10 mg/L UV
 pH= 7



(2-Theta-Scale)



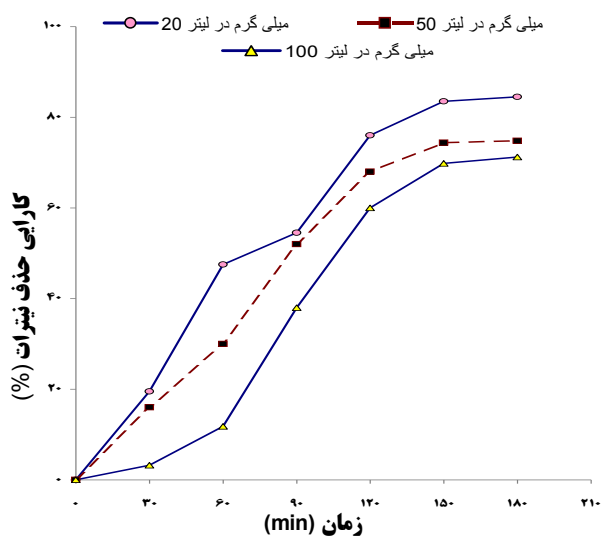
$C_0 = \dots \text{mg/L}$ pH
 $t = \dots \text{min}$ Ag-TiO₂ = \dots g/L : Ag-TiO₂/UV



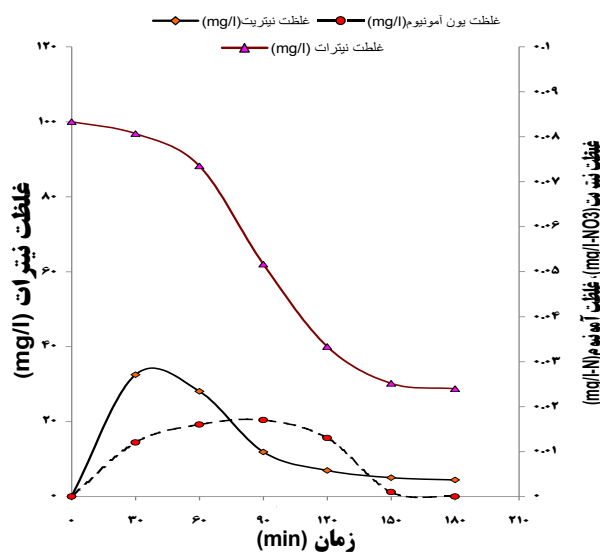
Ag-TiO₂/UV
 pH = \dots ; $C_0 = \dots \text{mg/L}$

Ag-doped TiO₂ Ti Ag
 " " " "

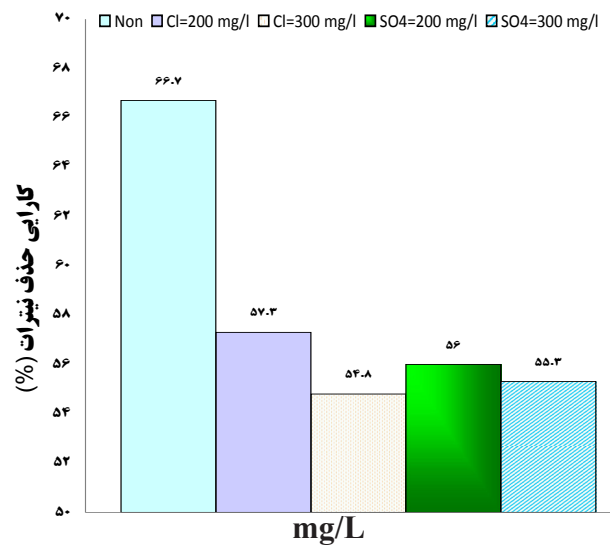
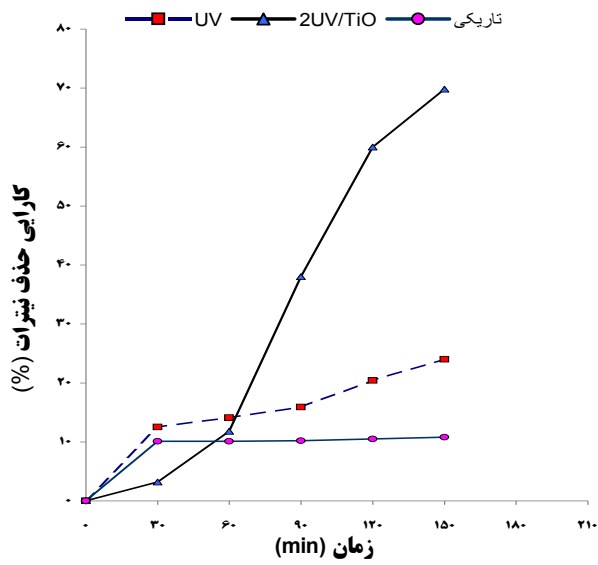
BET XRD TEM SEM-EDX
 Ag-doped TiO₂
 TEM
 / nm Ag-doped TiO₂
 TiO₂
 SEM-EDX / nm nm
 Ti Ag (at%L) flwt%L
 " " " "



Ag-TiO₂ = \dots g/L ; pH = \dots : Ag-TiO₂/UV

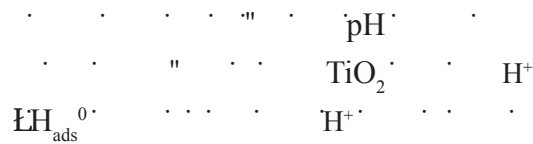


Ag-TiO₂ = \dots g/L , pH = \dots ; $C_0 = \dots \text{mg/L}$: Ag-TiO₂/UV



UV Ag-TiO₂/UV
 Ag-TiO₂ = 5 g/L, pH = 7, C₀ = 50 mg/L

t = 120 min
 Ag-TiO₂ = 5 g/L, pH = 7, C₀ = 50 mg/L



pH

pH

Ranjit

M-TiO₂

(5 g/L)

Ag-TiO₂

pH

TiO₂ P25

XRD

fL

Ag-TiO₂

BET

TiO₂

/ m²/g

fL ± m²/g P25

Ag-TiO₂/UV

pH

fL

pH

pH =

pH

pH

pH fL

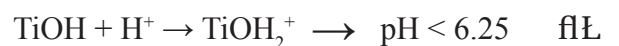
pH fL

pH fL

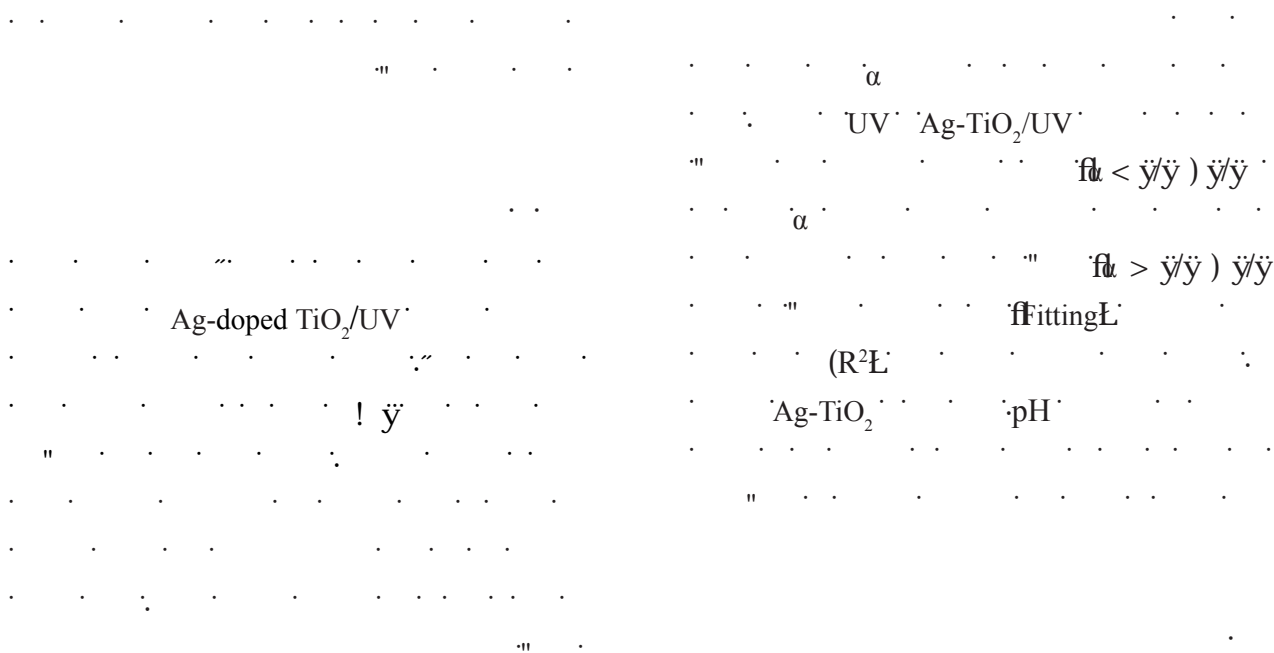
pH

fTiOH

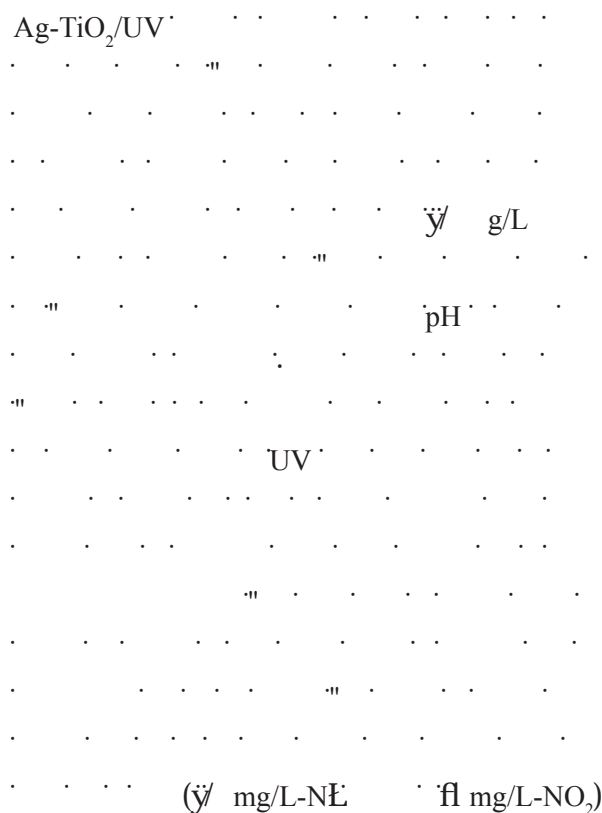
TiO₂



pH / g/L "fl - E
" " y/ g/L
" "
- E
Ag-TiO₂ fl
fl E " " " Yang "
Paracetamol
TiO₂
E
"fl
" "
y
"fl E Guo
TiO₂
TiO₂
nm
fl y Ranjit " "fl E
" " y/ g/L
n
" "
w pH E
fl E " fl
nm
E y Damodar .(y)
fl
" "
Ag-TiO₂
i ymin
" " ymin fl E
"
" "
" "
(E " "



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Photocatalytic Reduction of Nitrate in Aqueous Solutions using Ag-doped TiO₂/UV Process

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ABSTRACT

Background and Objectives: Pollution of water resources to nitrate is an environmental problem in many parts of the world. This problem possibly causes diseases such as methemoglobinemia, lymphatic system cancer and Leukemia. Hence, nitrate control and removal from water resources is necessary. Considering that application of nanomaterials in treatment of environmental pollutants has become an interesting method, in this research use of Ag-doped TiO₂ nanoparticles synthesized through photodeposition produced under UV irradiation was studied for removal of nitrate from aqueous solutions.

Materials and Methods: Three nitrate concentrations of 20, 50, and 100 mg/L were considered. In order to determine the effect of Ag-doped TiO₂ nanoparticles on nitrate removal, dosages of 0.1, 0.4, 0.8 and 1.2 g/L nanoparticles were used; pH range of 5-9 was also considered. The effect of Ag-doped TiO₂ nanoparticles both in darkness and under UV irradiation was studied. Moreover, the presence of chloride and sulfate anions on the system removal efficiency was investigated.

Results: The optimum performance of nitrate removal (95.5%) was obtained using nitrate concentration of 100 mg/L, in acidic pH and 0.8 g/L Ag-TiO₂. Increase of nanoparticle dosage up to 0.8 g/L, increased the removal efficiency, but for 1.2 g/L dosage of nanoparticles, the removal efficiency decreased. Maximum reduction performance without nanoparticles, under UV irradiation and under darkness conditions were 32% and 23.3% , respectively. In addition, we found that presence of sulfate and chloride anions in aqueous solution reduced efficiency of nitrate removal.

Conclusion: Results of this study showed that Ag-doped TiO₂ nanoparticles may be efficiently used for nitrate removal from aqueous solutions.

Keywords: Photocatalytic reduction, Ag-doped TiO₂, Nitrate, Aqueous solutions

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